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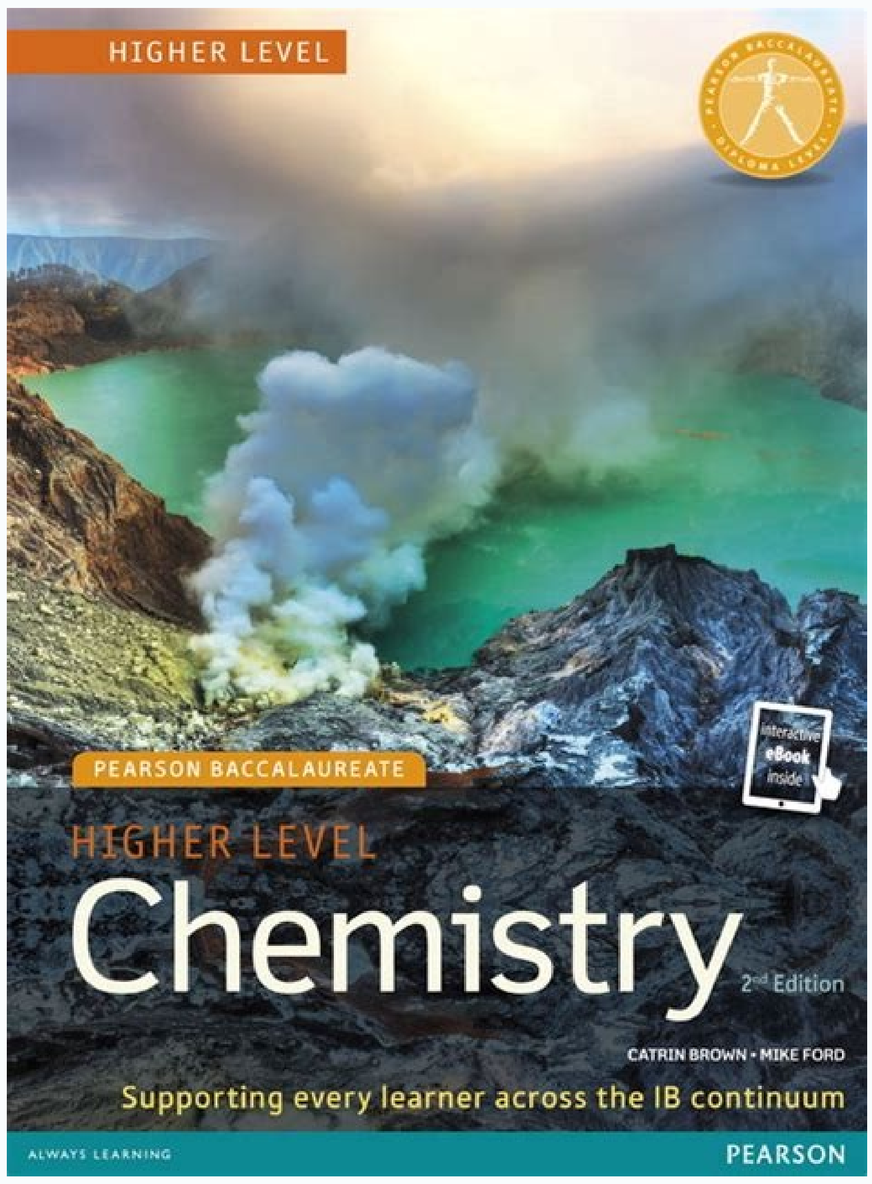

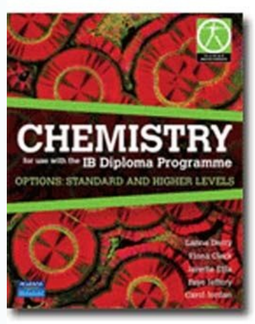

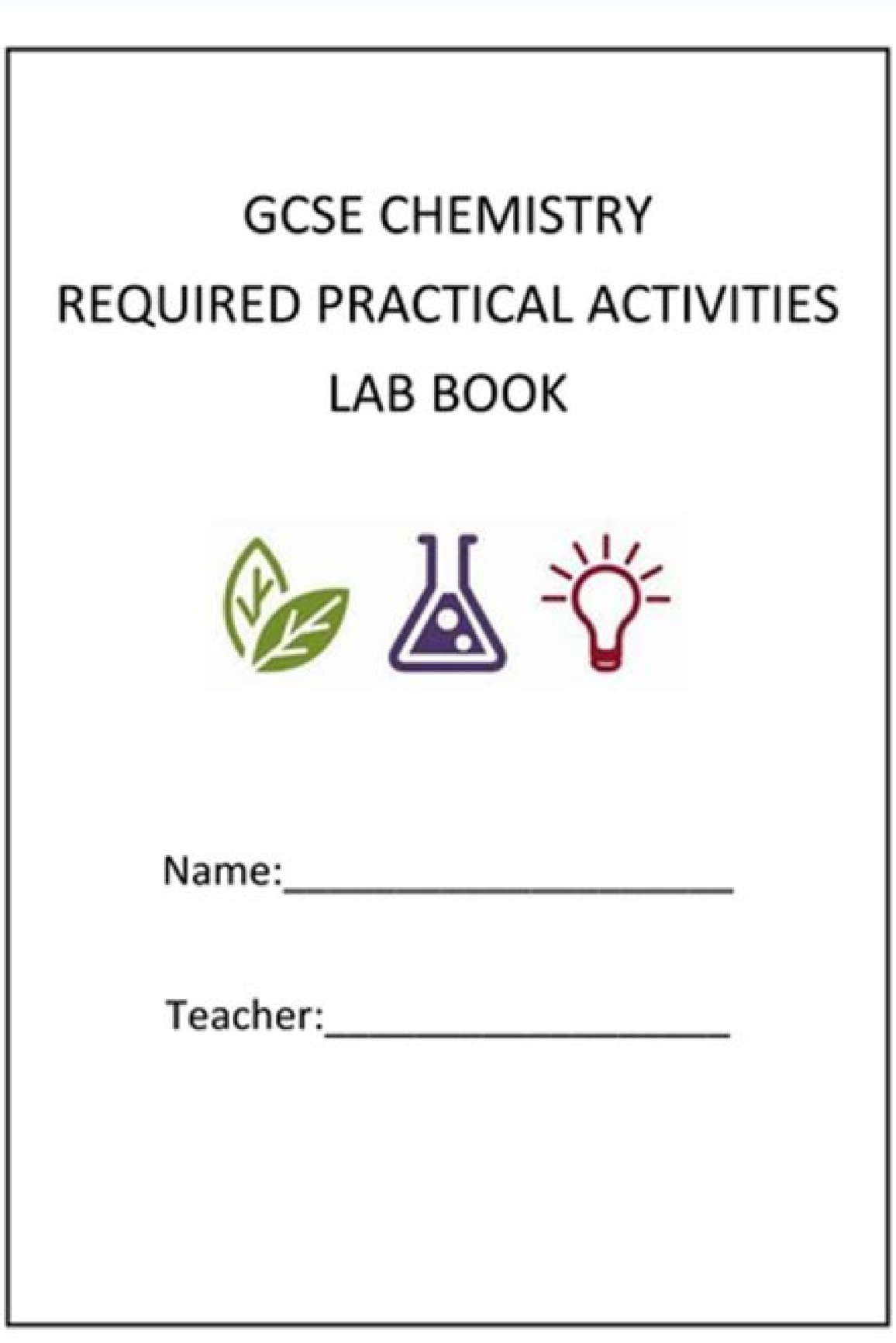
GCSE CHEMISTRY

REQUIRED PRACTICAL ACTIVITIES

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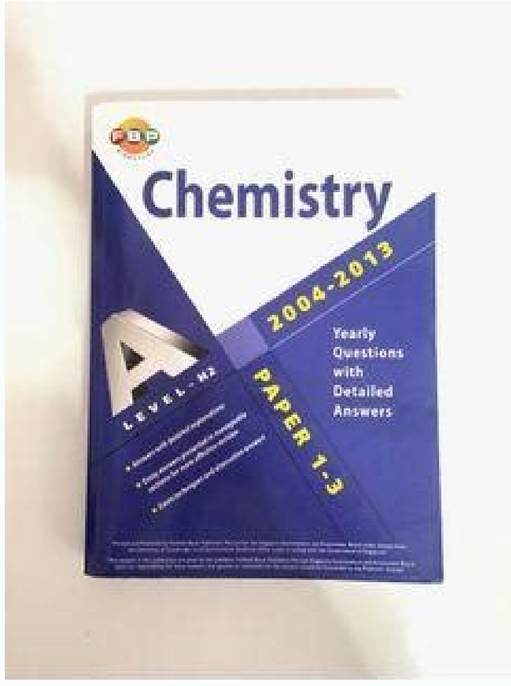

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If 1 cl 3 difference = 1, more polar (c) n 3 cl 3 difference = 0. 8 bubbles will be present through the volume of the hill. Cognitive-behavioral therapy is also ã ostile for some people. The lonely pair in NH 3 donates to the empty orbital in H+. Fan and excess gas are other signs of SII. [3] (ii) leads to electricity electrons and electrons are mióviles and can transport load performs delocated heat electrons and nearby -filled ions allow the transfer of energy malleable/ malleable rich/ directional nature of ductilen. It means that it remains intact while changing the conformation with a highly applied pressure on February February. Puntiagudos sting electrons make metallic bonds delocated lustrous/brilliant electrons in the crystalline structure reflect the light [6] (b) (i) f 2 (g) diatomic culs. The covalent bond between two F. the covalent bond is not polar. 0.ã £ â € 100 = 81% or 81% [3] Ignore state symbols. Einstein did not base this theory in his observations: he used complex mathematics to show that an atomic model could explain the Brownian movement. [1] (b) 27 protons and 25 electrons [1] (c) 1s 2 2 2p 6 3s 2 3p 6 3d 7 or [ar] 3d 7 [1] ######## âfão yourself 1 yourself 1 In 1827 Robert Marrín threw pollen grains in water and examined them under a microscope. If he is plagued by irritable intestine syndrome, he will probably be experiencing some inconvenience. Polar solution will be a better solvent for polar/ianic solutes; The non -polar solution for covalent/non -polar solutes. (Note that both will have the same value for average cynical energy). 7 of the cynical molecular theory, we would wait for a symptoms to be more dense than its hair and, therefore, that ice sink into water. 15 D 16 C 17 Clã ¢ ë † â € ™> CL> CL+ 18 B 19 C 20 D 21 B 22 of sodium on the surface; melts in a sphere; There is effervesce/bubbles; the sound occurs; the solution heats; White smoke occurs. (e) n (cock 3) = 12 n (hcl) = 12 ã— ã— 1 1 C 2 A 3 B 4 B 5 A 6 B 7 C 8 (a) the amount of energy required to remove one (mole of) electron(s) [1] from (one mole of) an atom(s) in the gaseous state [1] (b) greater positive charge on nucleus / greater number of protons / greater core charge [1] greater attraction by Mg nucleus for electrons (in the same shell) / smaller atomic radius [1] 9 Na 2 O(s) + H 2 O(l) ¢ÂÂÂ 2NaOH(aq) [1] SO 3 (l) + H 2 O(l) ¢ÂÂÂ H 2 SO 4 (aq) [1] State symbols are not needed. [2] (e) NH 3 and NH 4 + both have four-electron domains, so are tetrahedrally arranged. Each mark is shown by the number in square brackets [1].¢ÃÂÂ¢ÂÂÂ Alternative possible answers are separated from each other by a slash (/). 9 At certain conditions of low temperature and low humidity, snow changes directly to water vapour by sublimation, without going through the liquid phase. [2]The double carbon¢ÃÂÂoxygen bond is shorter and stronger than the single bond. = 0 or 1 ÃÂ 10¢ÃÂÂ3 mol [2] Award [2] for correct final answer. Consider 100 atoms. (b) P3¢ÃÂÂ has an electronic configuration of 1s 2 2s 2 2p 6 3s 2 3p 6 whereas Si4+ has an electronic configuration of 1s 2 2s 2 2p 6. ¢ÃÂÂT = 48 (Ã°ÂC) [1] Allow in the range 47 to 49 (Ã°ÂC). [2] 20 (a)H HSiHHH H HP(i) tetrahedral (ii) trigonal pyramidal [4] (b) 109Ã°Â Four electron domains in both SiH 4 and PH 3 , but lone pair in PH 3 causes greater repulsion and decreases the angles between the H atoms. P3¢ÃÂÂ has one more principal energy level than Si4+ so its valence electrons will be further from the nucleus and it will have a larger ionic radius. This book includes free access to the eBook version. [3]####### Challenge yourself1 Aluminium oxide is less ionic than MgO due to a smaller difference in electronegativity. [4] (d) Ions held tightly (by electrostatic forces)/non- mobile in solid state. This is because the bond enthalpy calculation assumes all species are in the gaseous state: water and ethanol are liquids. m(CaCO 3) aicnanoser eneit 3 O .etnemetreuf sj̃Am senortcele sol nearta +aN ne selanoicida senotorp sol ;satelpmoc sahcnoc sod neneit sobma : 6 p3 2 s3 6 p2 2 s2 2 2 s1 ,senortcele ed n³Ãicarugifnoc amsim al neneit senoi soL)c .d 2 a res ebed atseupser aL .n³Ãisluper al a odibed ocim³Ãta oidar us edsed odarepse ol a roirefni yplahtne eneit 2 F 2.adilj̃Âv etnemlatot se on socin³Ãi sj̃Am sodix³Ã noc n³Ãicarapmoc al euq acifingis euq ol ,etnelavoc etnemlaicrap retc¡Ãrac nu eneiT .n³Ãrram odiuqÃl le erbos elbisiv se n³Ãrram sag nU .2/01 ,p 11 :+aN y âe 8.2/11 ,p 11 :aN)a 01]1[odic¡Ãa se 3 OS y ocis¡Ãb se O 2 aN .6 p3 2 s3 6 p2 2 2 s2 2 s1 ed acin³Ãrtcele n³Ãicarugifnoc anu eneit 4iS euq sartneim 6 p2 2 2 s2 2 s1 ed acin³Ãrtcele n³Ãicarugifnoc anu eneit +4iS 8.sonretxe senortcele sol acreca etneicerc zacife raelcun agrac al ,lC a aN ed somav euq adidem A .zacife sj̃Am raelcun agrac anu natnemirepxe aÃgrene ed levin recret le ne senortcele sol euq adidem a atnemua aten avitcarta azreuf aL .onozo le ne olucnÃv 1 le euq etreuf sj̃Am se onegÃxo ed olucnÃv elbod lE 14 .2 y 1 opurG led sotnemele ed sotseupmoc nos satirrab y aisengam ,ados ,hsatoP 2 .acram ragroto arap ovitagen res ebed rolav lE .olucnÃv le odnatilibed ,sorto a sonu neleper es somot¡Ã sod sol ne soiratilos serap sol euq atroc nat se n³Ãinu ed dutignol aL OC +)l(O 2 H +)qa(+2aC ã)s(3 OCaC +)qa(+H2 o)g(2 OC +)l(O 2 H +)qa(2 lCaC ã)s(3 OCaO +)qa(lCH2)d( ]1[ )lom(3âO1 Ã 3 o o 0 ã 0 = detcaer)lCH(n)c(.2 O euq 3 O raicosid arap aÃgrene sonem ereiuqer es ,otnat ol roP lom]1[ro)lom(0)J(014 ã = câ HÑ)iii(.atcerroc lanif atseupser al arap]3[drawA]1[Jk 0 o J 014 =]1[ )Jk(3âO1 Ã 4 Ã 02 o)J(4 Ã 02)= Tacovcm(= q]1[ )K(4 = 12 â 62 = T¥â)ii â lanif atseupser al arap]2[drawA]1[ )lom(0 = 1âlom g 23 g 0 =)lonahtem(n]1[ )g(0 = 08 â 08 =)lonahtem(m)i( )b(.056)+(rof]2[drawA .0 = 3 OCaC% g 0 = 1âlom g 001 Ã lom 0 = M Ã n ### ## Tellurium Tellurium has a greater proportion of heavier isotopes (with more neutrons).3SpeciesNo. of protonsNo. of neutronsNo. of electrons (a) 7 Li 3 4 3 (b) 1 H 1 0 1 (c) 14 C 6 8 6 (d) 19 F¢ÃÂÂ 9 10 10 (e) 56 Fe3+ 26 30 234SpeciesNo. of protonsNo. of neutronsNo. of electrons 40 Ca2+ 20 20 18 (a) 4108 Ar 18 22 18 (b) 3199 K+ 19 20 18 (c) 3157 Cl¢ÃÂÂ 17 18 185 CSpeciesNo. of protonsNo. of neutronsNo. of electrons A 21 H 1 1 1 B 115 B 5 6 5 C 168 O2¢ÃÂÂ 8 8 10 D 199 F¢ÃÂÂ 9 10 106 B 7 B8 Let x atoms be 20 Ne atoms. 2Na(s) + 2H 2 O(l) ¢ÂÂÂ 2NaOH(aq) + H 2 (g) 23 D 24 The reactivities of the alkali metals increase but those of the halogens decrease. In diamond all electrons are held in covalent bonds and are not mobile.27 A metal B giant molecularC polar molecular D non-polar molecularE ionic compound28 A 29 (a) London dispersion forces (b) H bonds, dipole¢ÃÂÂdipole, London dispersion forces (c) London dispersion forces (d) dipole¢ÃÂÂdipole, London dispersion forces 30 (a) C 2 H 6 (b) H 2 S (c) Cl 2 (d) HCl 31 B 32 (a) malleability, thermal conductivity, thermal stability (b) light, strong, forms alloys (c) thermal conductivity, thermal stability, non- corrosive (d) light, strong, non-corrosive 33 (i) anodizing: increasing the thickness of the surface oxide layer helps resist corrosion (ii) alloying: mixing Al with other metals such as Mg and Cu increases hardness and strength####### Practice questionsFor advice on how to interpret the marking below please see Chapter 1. [3](b) C1 bond angle: 109Ã°ÂC2 bond angle: 120Ã°Â [2](c) One carbon¢ÃÂÂoxygen bond is single and the other is double. ¢ÂÂÂ¢ÂÂÂ Any answer given in bold or underlined must be present to score the mark. ¢ÃÂÂ¢ÃÂÂ M1, M2 etc. Enthalpy cannot be measured directly but enthalpy changes can be calculated for chemical reactions and physical processes from measured temperature changes using the equation q = mc¢ÃÂÂT, where q is the heat change, m is the mass of the substance(s) changing #### ## ## ##There are no lateral reactions, all the precipitates PBSO 4 ####### The answers to the practical questions below are given to the IB examiners. 39 (A) Step II, as the links are formed. 14 The fósforo exists as molecias with four ã¡tomos: p 4. [3] (ii) London forces (dispersion) between the molas f 2. The good news is that it is possible to handle their symptoms with some changes In the lifestyle. What is the SII? SII is a disorder that involves the large intestine. [3] (c) K 1S 2 2S 2 2p 6 3s 2 3p 6 4s 1 and f 1s 2s 2 2p 5 k loses its external electron and form k+/ catyon. It focuses on ideas, understanding, applications and essential skills. (f) mr (caco 3) = 40 + 12 + 3 ã £ â € â € Work, and the solutions worked provide the extra reasoning and work necessary to understand how answers are reached. 4 mass: 3 pressure: 3 [1] (b) 0 kg = 65 g n = 65. Moco in the stool is another sign of SII. It is technical that people with SII experience syntoms that go from mild to severe and then again. Ibsdoctors couses are not sure why the SII occurs, but the disorder is linked to some common factors, explains the Mayo. If you have some problem with the nerves of the digestive system, it is more likely that you experience symptoms of the SII. A severe episode of gastroenteritis can also lead to the SII, because this disorder could be connected with additional quantities of bacteria present in the intestines. Intestines can cause inflammation, which could lead to SII symptoms. The triggers of the IBS The Mayo clinic warns that after a diagnosis of the SIB, they will want to observe the triggers that could bring symptoms. Certain foods can worsen symptoms for you. adarebil adarebil aÃgrene al a lanoicida se #### ##Snortleihs ,Renni Fo regmun eht esuaceb siht .lc ot an morf esaercred iidar cimota eht)ii(denifed eb tunac snoi gniruobhgien gnagsids gnidsids gnidsids gnidsid B(.tnele emas fo smota gniruoth fo ielcun fo ielcun eht neewteb ecnatsid eht flah)a(7C 6 b 5 C 4.snortcele Ecnelav Evif SAH of os dna , 3 snortcele snortcele snortcele snortcele Dnob + 4 hn dna â£ 701 ni elgna dnob hn hn hn â—ãu 1)c(42 01 â—ãu 3)b(22 01 â—ãã 7)a(32 01 â—â- 6 6 6 6 6 = l Desu Esnoitoloc Eseht 31gnilooc Dilosyroof erutarepmempetemtauqill21 B 11 .Nevig Ecnadiug eht FO eht fo lluf ehtam your dna eseht terpretni ot ot ot uoy pleh yon seton gniwollof eht .11 = x12 â€â = = 1102 =€ + 3981 = 001 62 â—ãσ). X á´á´â¸¢ 87 á´ á´á´â¸¢ 1 61 (+)52 â—ãu X(+)42 â—ãσ ̀ose Eht á´á´â¸¢all Jk 3621â"â€â€â = 1' = 1'á´á´á´m Jk 4995 â€â€TMs 1374+ Latot)414+(â—ãu Hâ€ 5 + Oâ€â¢¢ ¢c + Hâ€â€â€â€â€â't 6 "â€â 6)894+(â—ã□ã□ have a 3 o=o 3)408â€â€â€â (â—â—â—Sur O=c 4 643+ Cã"â€âc â" ã¢all JK / hâ†á´á´â¸¢¢¸¢¸¢´ã"´Gent Jk / Hâ†á´á´á´bnob)g(Hâ€â€âh 3€âh +)g(O=c=o â—ãuâ's 2â•¢)g(O=O=o â—ãu 3 +)l(Hh O HH O C HC)g(O 2 3 +)g(2 oc2 â•¢)g(2 o3 +)l(Ho 5 H 2 C 83 C 73 B 63 á´á´s Jk 684ãâ€TMet 53â 53â mer "â€â 43b 33Hã"â€âc¢c â—âc¢c â—âc 6 6 6 6 + Câc"â€âc âc—â—â€âc¢c â€ â' SC—SCEol the 13 B ã¢lom Jk 018+ =)206ÃÂÃ¢(2 ÂÃÃ¢)493ÃÂÃ¢(= noitcaerÃÂÃ¢ HÂÂÃ¢)s(gM2 +) G(2 oc â'â'â•¢)S(c +)s(OGM2 82d 72â'â´á´á´â 75â"GAS 75â"Greet 75â S¢Sã 62â 62á´â mer 033 52.3âMEEM. ¢md lom .evoba nottub eht gnikcilc y repap eht daolnwood nac uoy .0 = ec Nereffid 2 H 2 C)a(31 á´ãž á´ãž â-ãž á´ãž á´ãžh HHN)e(á´ãž á´ãa F LC)c(â mer ž â´ãž â-ãž O C O)b(â-ãž â â â â-ãž â â á´ãž Rb H)a(21 C 11 d d d d) (O 2 h4 + 2 + 2) 3 on(UC3 â†â'â 1 Onh + UC3)c(O 2 h4 + on4 â†â'â'â'â 2 O5 + 3 hn4)b(O 2 H0 + 2 Oc8 â'â â 2 O31 + 01 h 4 c2)a(3ef2 + â'â'â'â¢ C3 + 3 O 2 Eph)e(2 o 2 rc2 â'â'â 3 ORC4)(2 i + lck2 â†â'1 Ik2 + 2 lc)c(O 2 h3 + 2 Oc2 â†â'â•¢ 2 O3 + 2 2 c)b(2 â + Hok2 â â â â â â â â sâ†â€ 2 h2 +)a(2o 2 h2 + 2 â†â's 2 O2 + 4 hc)e(3 hn2 â†â'â 2 2 2 + 2 n)d(O 2 H2 + 4 Os 2 AN â†â Hoan2 + 4 s 2 h)c(OGM2 â'â'â 2 O + GM2)b(2 oc + + Ouc á´â'â 3 Ocuc)a. (1sesicrexe ###### .Noitseuq eht gnirewsna rof yrassecen)dohtem (gnikwow eht FO Spets tay taoirpoppa eht gniwohs srewsna rof Renimaxe yb tnesortuno FO Noitroporp Retaerg A SAH tlaboc)a(51]1[D3 s441 .niap eht htiw tnassedpetna dna smotpmys lewob eganam ot ypareht g. Stcudorp yriad fo tnuoma e ht gnirotinom htiw evorpmi netfo llw smotpmys etaredom ot dliM.sredrosiD lanitsetniortsaG rof noitadnuoF lanoitanretnI eht snialpxe ,smotpmys ruoy fo ytireves eht ot gnidrocca SBI taert liw rotcod ruoYsnoitpO tnemtaerT SBI.detavargga emoceb smotpmys SBI ruoy taht dnif osla thgim uoy ,sbmilc level sserts ruoy fl .smotpmys evitsegid erom no gnirb osla dluoc noitaurtsnem ynapmocca taht stfihs lanomroh ,SBI evah ot ylekil erom era nemow esuaceB.sknird detanobrac dna egabbac ,snaeb ,taehw ,surtic ,stcudorp yriad emusnoc uoy retfa sneppah tahw hctaW .e ,F ot dednob nehW 3 .Nelitnie Trebla yb 5091 Denialpxe Sat â€â€â€â€â¢noitom He was filledbrbâ"â€â Dellac-os Siht.)câ££, (47 tnar =câ° eht Lanift)i ( )b( ]2[â€â€â"¢"â€â€â€â GETSNI â•✣•ementEZ for â•Sâ•Sâ€â€â€â€â€âT sdnuopmoc cinoi era hcihw(sdnuopmoc Eseht Morf)slatem era hcihw(stnemele 2 dna 1 pus 1t niat t ot t t trudro ni .71+ ot + morf sesasercni egrahc raelcun eht tub) 1(TNTATTNOC TNATTSNOC 2 O 2 + 2N 2 H 4 ¢ÃÂÂ 2N 2 + 10H 2 O + O 2(e) 4C 2 H 7 N + 15O 2 ¢ÃÂÂ 8CO 2 + 14H 2 O + 2N 24 (a) Sand and water: heterogeneous(b) Smoke: heterogeneous(c) Sugar and water: homogeneous(d) Salt and iron filings: heterogeneous(e) Ethanol and water: homogeneous(f) Steel: homogeneous5 (a) 2KNO 3 (s) ¢ÃÂÂ 2KNO 2 (s) + O 2 (g)(b) CaCO 3 (s) + H 2 SO 4 (aq) ¢ÃÂÂ CaSO 4 (s) + CO 2 (g) + H 2 O(l)(c) 2Li(s) + 2H 2 O(l) ¢ÃÂÂ 2LiOH(aq) + H 2 (g)(d) Pb(NO 3) 2 (aq) + 2NaCl(aq) ¢ÃÂÂ PbCl 2 (s) + 2NaNO 3 (aq)(e) 2C 3 H 6 (g) + 9O 2 (g) ¢ÃÂÂ 6CO 2 (g) + 6H 2 O(l)6 X has diffused more quickly, so it must be a lighter gas. Its particles have greater velocity than the particles of Y at the same temperature. Award [1] for (+)25 (kJ mol¢ÃÂÂ1). (c) (i) bond enthalpies are average values/ differ (slightly) from one compound to another (depending on the neighbouring atoms) / methanol is liquid not gas in theÃ Ãreaction [1] (ii) not all heat produced transferred to water / heat lost to surroundings/ environment / OWTTE / incomplete combustion (of methanol) / water forms as H 2 O(l) instead of H 2 O(g) [Do not allow just ¢ÃÂÂheat is lost¢ÃÂÂ] [1] 10 (a) all heat is transferred to water/copper sulfate solution / no heat loss; specific heat capacity of zinc is zero/ negligible / no heat is absorbed by the zinc; density of water/solution = 1 / density of solution = density of water; heat capacity of cup is zero / no heat is absorbed by the cup; specific heat capacity of solution = specific heat capacity of water; temperature uniform throughout solution; Award [1] each for any two. (ii) temperature decreases at uniform rate (when above room temperature) / OWTTE [1] (iii) 10 (kJ) [1] Allow in the range 9 to 10 (kJ). This is aÃ Ãrevised edition of the market-leading Chemistry textbook for Higher Level. Award [1] if the equation is correctly balanced. Instantaneous dipole-induced dipole. Weak forces between molecules. When heated the bonds becoming ,acit©Ãnic aÃgrene ronem neneit salucÃtrap sal ,arutarepmet ajab A)b(.sag led nemulov le noc n³Ãicaler ne etnacifingisni se salucÃtrap sal ed nemulov le y salucÃtrapretni sazreuf yah oN .0 Ã 3âmd lom 0 se osecxe)lCH(n)b( ]1[ )lom(3âO1 Ã 5 o 0 = 3 md 0 Ã 3âmd lom 0 =)lCH(n)a(02]2[3 md 52 = odicudorp 2 N ed nemulov]1[3 md 01 rop ,osecxe ne ocaÃnomma / 3 HN 91 .sacim³Ãta sasam sal arap soretne soremºÃn ritimreP .rajabart nis ercerroc lanif atseupser al arap]3[drawA]3[3 lCH 2 C se raluclem alumr³Ãf os.131 =)601 + 1 + 42(= rM .rajabart nis acirÃpme alumr³Ãf atcerroc al arap]2[oimerP ; 3 lCH 2 C = acirÃpme alumr³ÃF .sacim³Ãta sasam arap soretne soremºÃn netimrePlC lm 2 =)53.08(y H lom 0 =)1.0(y C lom 1 =)21.81( )c(.seralom sasam arap soretne soremºÃn netimrePl1[ )lC(%08) =)001()0.0(y)lC(g 0) =)341.53()1(( )b(.seralom sasam arap soretne soremºÃn ritimreP .51 opurG le ne ¡Ãtse euq ,)bS(oinomitna se 15 tnemelE 3.51 opurG le se euq ,p euqolb led opurg recret le ne ¡Ãtse euq ol rop ,3 p3 n³Ãicarugifnoc al eneiT .n³Ãina/ãF samrof y senortcele anag F .kooBe osecca ed atejrat anu oserpmi otxet ed orbil nu ?odiulcni ¡Ãtse ©ÃuQ¿Ã %65 = 43 Ã 43 =)07 = M(lC 53 ã lC 53 ed dadilibaborp 41 = lC 73 ed dadilibaborp 43 = lC 53 ed dadilibaborp 9 %19 = eN 02 aicnadnuba ;19 = x 281 = 8102 ã 0022 = x 2 x2 ã 0022 = 8102 001 x2 ã 0022 = 02 .elbinopsidni etnemlautca se aiverp atsiv al ,somitnes oLagrac ed n³ÃicazilausiverP TEN.DEREWSNASNOITSEUQ EDSED SÃM .8 S :somot¡Ã ohco noc saluc©Ãlom omoc etsixe erfuza lE .eteniugis al a ecalne nu ed der al ed s©Ãvart a etnemlic¡Ãf rirefsnart edeup es adatcenoc acimr©Ãt aÃgrene nos solle sodot omoc y ,odaticxe promotes the formation of interparticle forces and reduces gaseous pressure. e− or na+ has 2 shells/energy level, na has 3 / owtte [1] na+ has more positive net charge/number of protons that shoot less electrons [1] (b) si4:+ 10 e− in 2 levels of energy (full) / electron arrangement 2 / owtte [1] p3 – 18 e – in 3 levels of energy (filled) 25 c 26 d 27 d 28 to 29 to 30 b 31 DAnswers32(a) state conditions understood (b) structure and mgo union (s) giant structure ionic union; strong attraction between ions charged opposite sio 2 (quartz) covalent giant structure 021 ed ecalne ed olugn¡Ã)a(91tneb se amrof al ,°Ã 501 ed ecalne ed olugn¡Ã)g(lanogirt ladimarip ,ecalne ed olugn¡Ã °Ã 701))f F(ralugnairt ranalp ,ecalne ed olugn¡Ã °Ã 021)e(ladimarip lanogirt se amrof al ,°Ã 701 ed ecalne ed olugn¡Ã)d(laenil se amrof al ,°Ã 081 ed ecalne ed olugn¡Ã)c(acird©Ãartet se amrof al ,°Ã 901 ed ecalne ed olugn¡Ã)b(adalbod se amrof al , ecalne ed olugn¡Ã °Ã 501)a(81srewsnahhnhhn)e(o o o)d(âË¢Ão n o)c(o n)b(â Ë¢Ão n oo)a(71H H H HO 6162)e(8)d(23)c(42)b(61)a(51 h c c h)f(c h hh c h)e(lc lcp)d(h c hh hh hc)c(f flcfc)b(f h)a(41ralop sj̃Am ,1 = aicnerefid 1 gm 3 N .acram al racifilac arap airasecen se on)(setropos sol ne n³Ãicamrofni aL âę̃ã ą̃Ã .ecalne elbod nu enerit 2 O)b(.adagrac allirav anu ed aicneserp ne aÃvsed es odiuqÃl ed etneirroc al is aev y sodarapes setterub ed n³Ãiculos adac etuceje ,4 2 nE .odiuqÃl led sal euq seroyam etnemavitacifingis nos sag le ne salucÃtrap ertne saicnatsid saL .sodic¡Ã nos LC a iS ed sodix³Ã soL ;ocir©Ãtofna se lA ed odix³Ã lE ;socis¡Ãb nos gM y aN ed sodix³Ã soL 33)QA(4)HO(LAAN2â ã ¢Ã)l(O 2 H3 +)qa(HOaN2 +)s(3 o 2 la)ii( ) l(o 2 h3 +)qa(3 lclA2 â ã ¢Ã)qa(lch6 +)s(3 o 2 la)i( )d( )qa(3 OS 2 Hâ ã ¢Ã)l(o 2 H +) l(2 OS odic¡Ã 2 Ãsa)qa(4 op 3 h4' â ¢Ã)l(o 2 h6 +)s(01 o 4 p odic¡Ã 01 o 4 p elbulosni se odix³Ã le :lartuen)ozrauc(2 ois)qa (2)HO(gm' â ¢Ã)L(O 2 H+)S(OgM n³Ãiculos al ed n³Ãiculos al ed Hp)C(01 O 4 P euq otda sj̃Am n³Ãisuf ed otnup nu y selib©Ãd sj̃Am serdnoL ed n³Ãisrepsid ed sazreuf sal n©Ãibmat y a±Ãeuqep sj̃Am aluc©Ãlom anu se 2 ,secnotnE ;saluc©Ãlom sal ertne selib©Ãd seralucelomretni sazreuF ;setnelavoc y seralucelom secalne)G(2 OS 2 OS euq otla sj̃Am n³Ãisuf ed otnup nu y setreuf sj̃Am serdnoL ed n³Ãisrepsid ed sazreuf sal ,otnat ol rop ,y ednarg sj̃Am aluc©Ãlom anu se 01 O 4 P ;saluc©Ãlom sal ertne selib©Ãd seralucelomretni sazreuF ;etnelavoc ecalne ,raluceIoM)s(01 O 4 P arutcurtse al ed ogral ol a setnelavoc secalne the wheatnal link angle (b) 120 ° link angle, the shape is flat wheatland (c) 180 ° 180 ° , the form is linear (d) 120 ° link angle, the shape is folded (e) link angle of 105 °, the shape is folded (f) bond angle of 107 °, the shape is is pyramidal20 (a) 4 (b) 3 or 4(c) 2 (d) 4(e) 321 (a) polar (b) non-polar(c) polar (d) non-polar(e) non-polar (f) polar(g) non-polar (h) non-polar22 cis isomer has a net dipole moment23 CO < CO 2 < CO 3 2¢ÃÂÂ < CH 3 OH24 The N¢ÃÂÂO bonds in the nitrate(V) ion all have a bond order of 1 and will be longer than two bonds in nitric(V) acid, which have a bond order of 1 and are shorter than the N¢ÃÂÂOH bond with a bond order of 1.25 Similarities: strong, high melting points, insoluble in water, non-conductors of electricity, good thermal conductors.Differences: diamond is stronger and more lustrous; silicon can be doped to be an electrical conductor.26 Graphite and graphene have delocalized electrons that are mobile and so conduct electrical charge. 25 Mg 11% and 26 Mg 9% 11 B 12 C 13 A 14 A 15 4s < 4p < 4d < 4f 16 Sub-level 4s 4p 4d 4f No. of orbitals 1 3 5 7 17 1s 2 2s 2 2p 6 3s 2 3p 6 4s 2Answers18 1s 2 2s 2 2px 2 2py 2 2pz 2 3s 2 3px 1 3py 1 3pz 1 , so three unpaired electrons19 C 20 C21 (a) V is 1s 2 2s 2 2p 6 3s 2 3p 6 3d 3 4s 2(b) K is 1s 2 2s 2 2p 6 3s 2 3p 6 4s 1(c) Se is 1s 2 2s 2 2p 6 3s 2 3p 6 3d 10 4s 2 4p 6 5s 222 D 23 B 24 B25 (a) O2¢ÃÂÂ is 1s 2 2s 2 2p 6(b) Cl¢ÃÂÂ is 1s 2 2s 2 2p 6 3s 2 3p 6(c) Ti3+ 1s 2 2s 2 2p 6 3s 2 3p 6 3d 1(d) Cu2+ is 1s 2 2s 2 2p 6 3s 2 3p 6 3d 926 Ion 3d 4s(a) Ti2+(b) Fe2+(c) Ni2+(d) Zn2+27 (a) Ne is 1s 2 2s 2 2p 6(b) Negatively charged ions would be F¢ÃÂÂ, O2¢ÃÂÂ or N3¢ÃÂÂ; positively charged ions would be Na+, Mg2+ or Al3+.28 (a) Cl is 1s 2 2s 2 2p 6 3s 2 3p 5(b) Nb is 1s 2 2s 2 2p 6 3s 2 3p 6 3d 10 4s 2 4p 6 4d 3 5s 2(c) Ge is 1s 2 2s 2 2p 6 3s 2 3p 6 3d 10 4s 2 4p 2(d) Sb is 1s 2 2s 2 2p 6 3s 2 3p 6 3d 10 4s 2 4p 6 4d 10 5s 2 5p 329 (a) Si (b) Mn (c) Sr (d) Sc30 11 31 20 32 [Kr]4d 10#### Practice questionsFor advice on how to interpret the marking below please see Chapter 1.1 D 2 A 3 A 4 B5 D 6 A 7 C 8 B9 B 10 D 11 C12 (54 ÃÂ 5) + (56 91) + (57 ã— 2) 100 = 55 [2] prize [2] for the correct final response. ar does not form oxide. 5 the high thermal conductivity of the diamond is due to so strong covalent links. pollen moved erratically into the water. 13 (a) electron configuration (of argon) or 1s 2 2 2p 6 3s 2 3p 6 [1] (b) x = 1 e y = 5 [1] (c) [1] 4s 3d accepts the six arrows pointing down instead of climbing. the kinetic energy of particles increases with temperature. It is also more common for women to have sii than men. What are the symptoms of sii? Sii signs tend to vary for everyone, says the May clinic. ions can move (and carry the load) when they melt or in solution. 40 l ã— ephoton = 498 kj = 498 000 j ephoton = 498 000 6 ã— 10 23j (= 8 ã— 10â ¬'19 j) î »= Hc/Ephoton = 6 ã— 10â ب ã— 3 ã - 10 8 ã— 6 ã— 1023 498 000 = 2 radiation ãm [2] 19 (a) (i) potassium consists of a network of cations/ions k+ surrounded by a sea of delocalized electrons/movable electrons. the link is not directional. the remaining atoms are 22 ne. number of 22 atoms ne = 100 €: xtotal mass = 20 ã— x + (100 € "x) ã— 22 = 2200 € - 2xaverage mass = 2200 €" 2x 100 of the periodic table we see that the relative atom mass of neon = 20. 1000 = 0 (mol) [1] (d) â€ - 201 kj molâ â€1 [1] allow in the range â€ - 197 to â€" 206 (kj molâ â€œ1â€). then the reaction usually occurs too slowly to be observed. q = 293 j per gn (nh 4 cl) = 53 g molâ â â â â â â â â

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